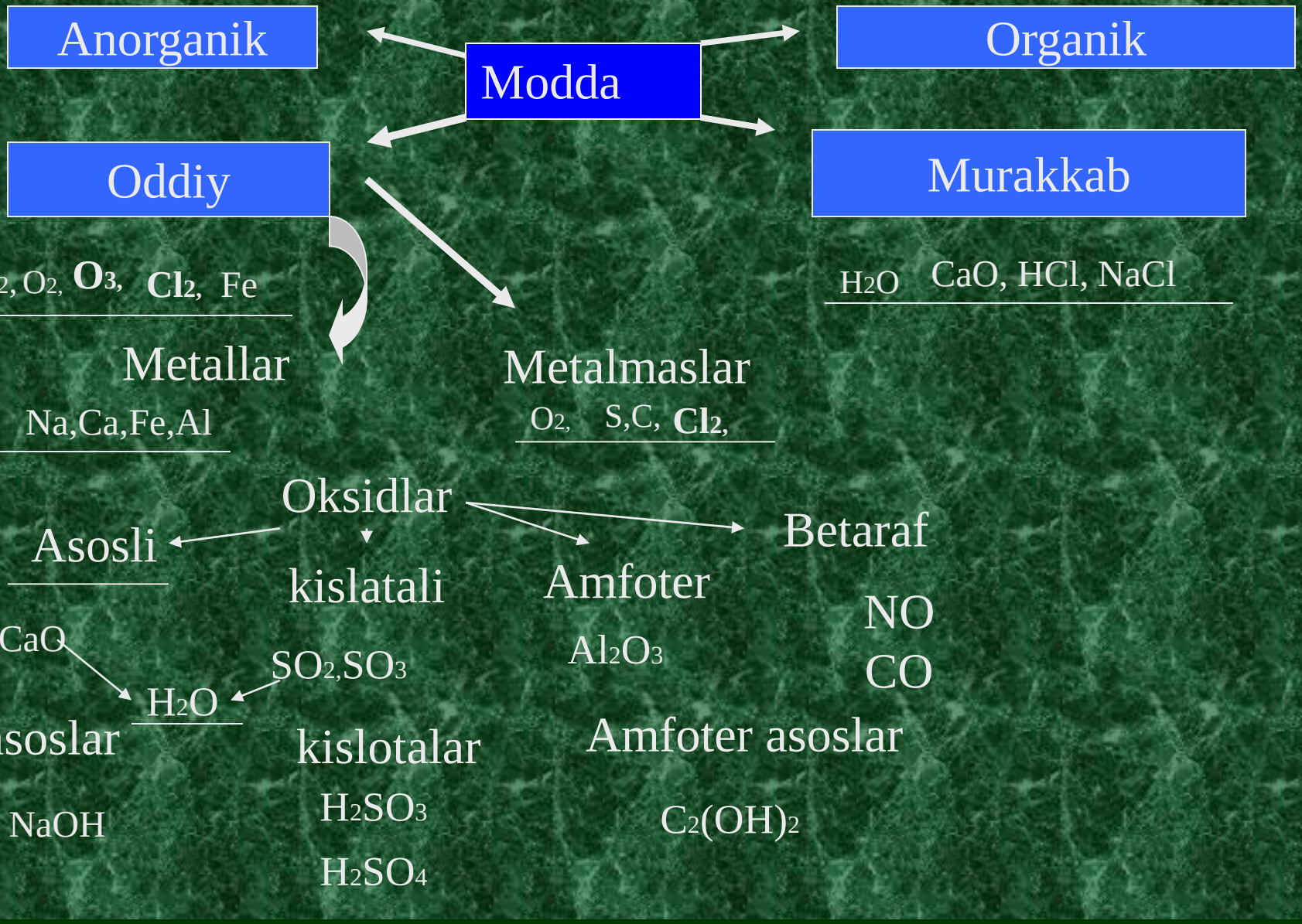
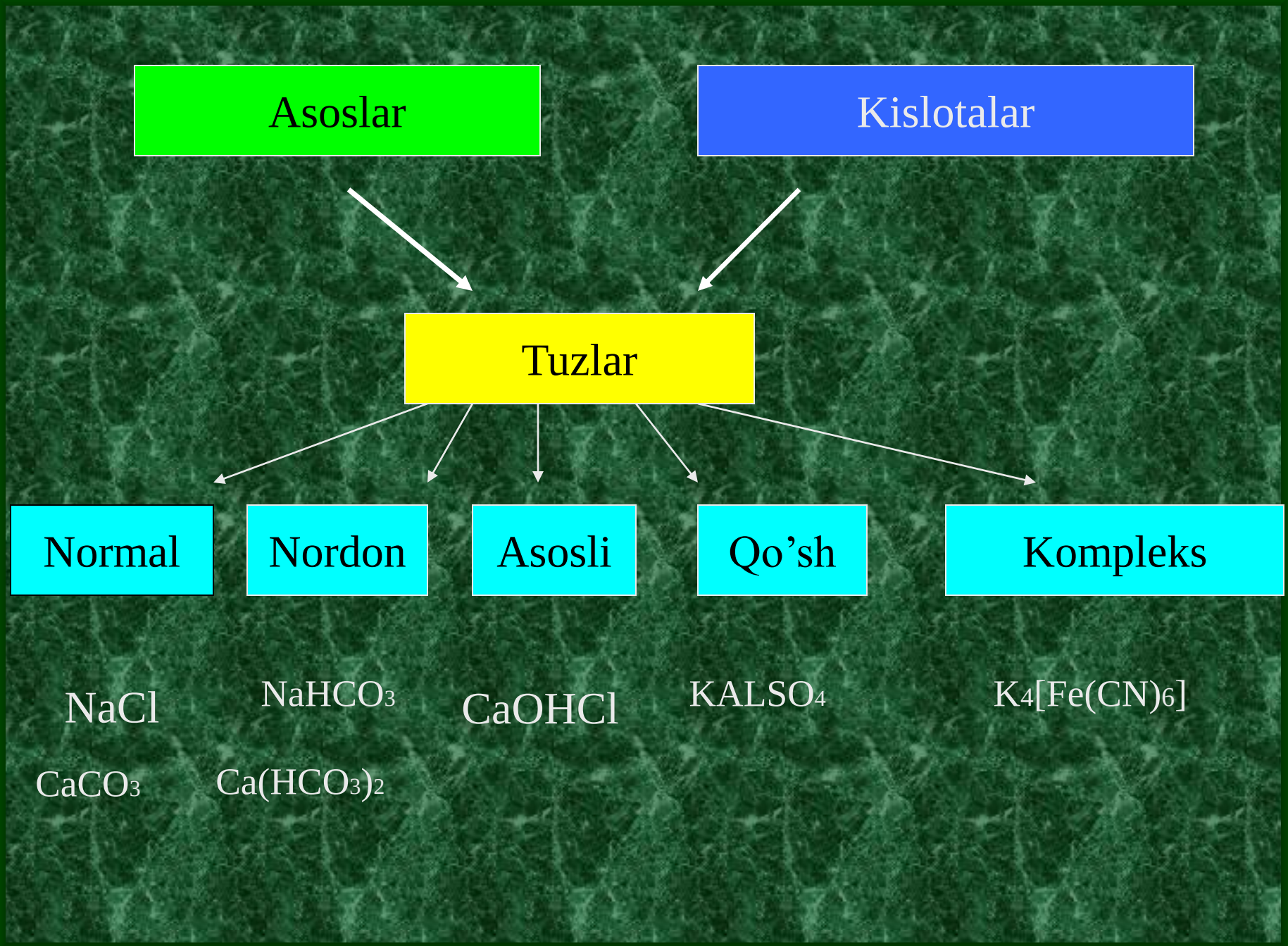
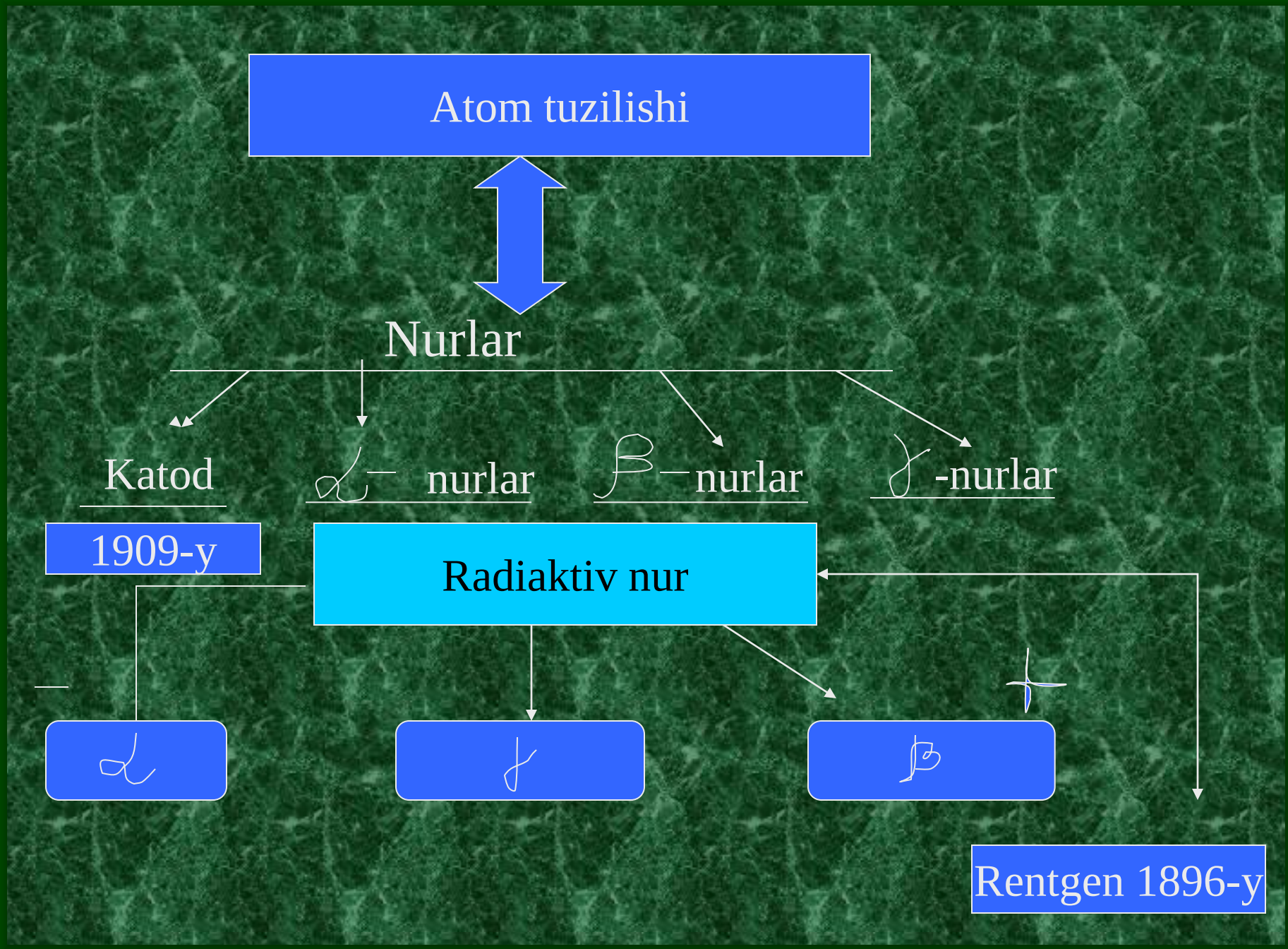
Anorganik birikmalarning sinflari




Asoslar
Kislotalar
Tuzlar
Normal
Nordon
Asosli
Qo’sh
Kompleks
NaCl
NaHCO3
CaOHCl
KALSO4
K4[Fe(CN)6]
CaCO3
Ca(HCO3)2


Atom tuzilishi
Nurlar
Katod
nurlar
nurlar
-nurlar
1909-y
Radiaktiv nur
Rentgen 1896-y

Radiaktiv yemirilish

1-qator	2-qator	3-qator
Ach=235	Ach=238	Ach=232
Aktiv uran	Uran qator	Toriy
$7\alpha\ 4\beta$	$8\alpha\ 6\beta$	$6\alpha\ 4\beta$
^{207}Pb	^{206}Pb	^{208}Pb

Atom yadrosini tuzilishi

Praton

Neytron

Elektron

N=A-Z

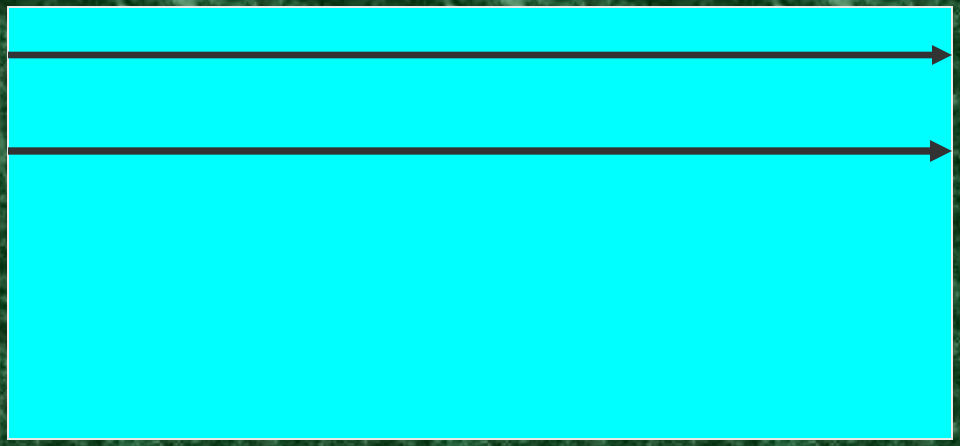
Yadro zaryadlari teng

Izatop
Yadro zaryadlari
Bir xil

Izobar
Yadro zaryadlari
turlicha

Izoton
O’zgarmas

D.I. Mendeleyev davriy sistemasi



Ionlanish energiyasi
E.M

Elektronga moyillik

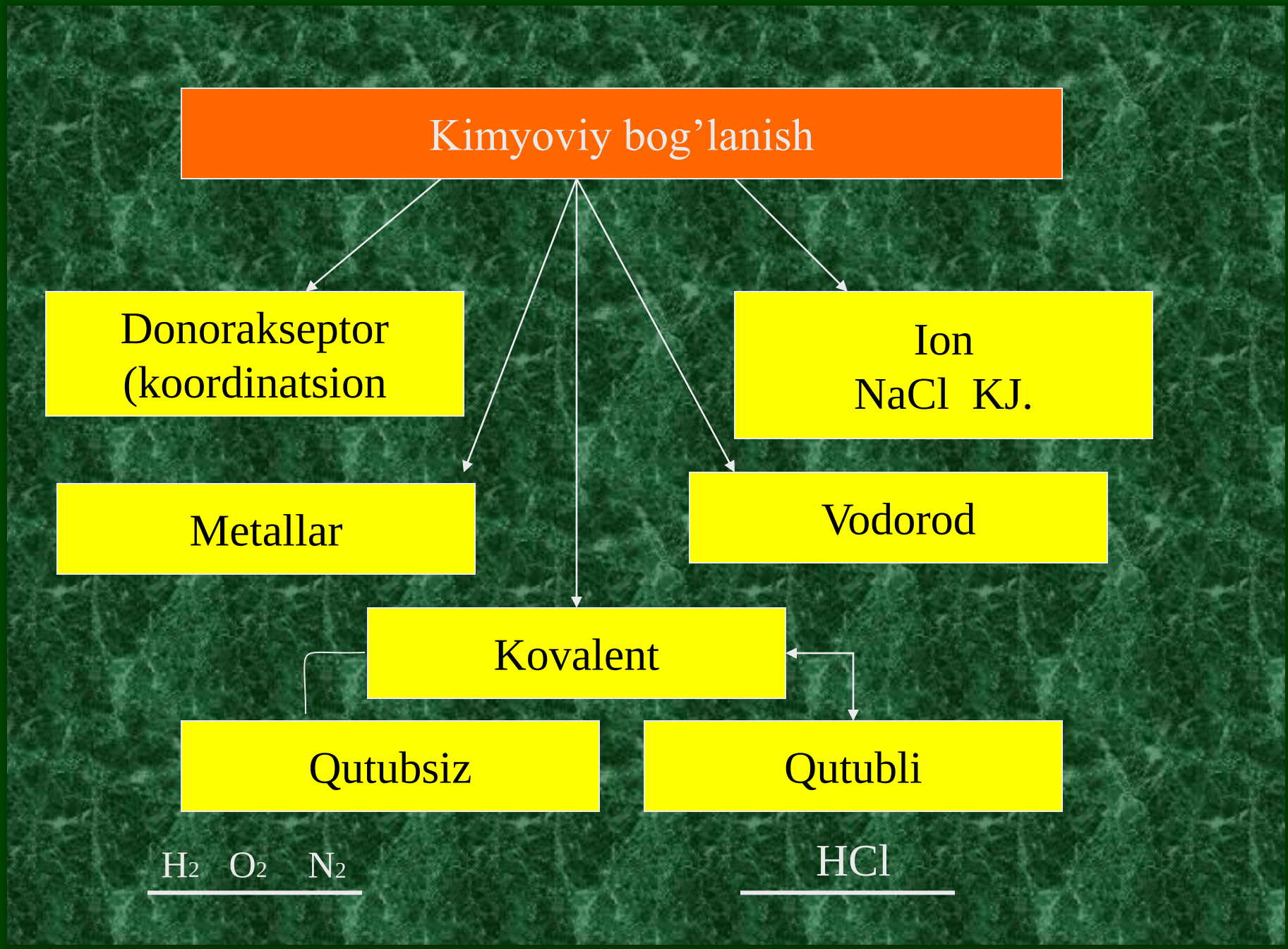
+

Ionlanish energiyasi

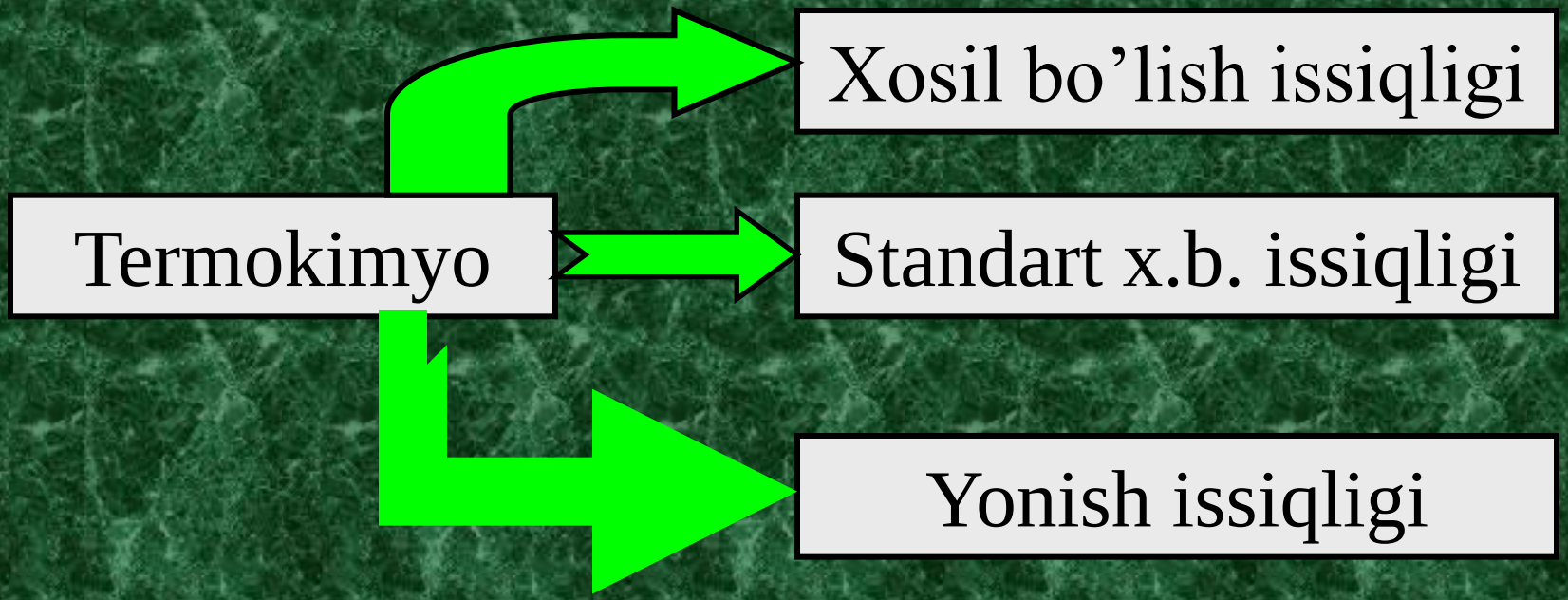
=

E.M


Kimyoviy reaksiya tezligiga ta’sir etuvchi omillar
Modda tabiati
Konsentratsiya
Katalizator
Xarorat
Kataliz
Aktivator
Ingibator
Gamogen
Geterogen
MnO2
Fe


Kimyoviy bog’lanish
Donorakseptor (koordinatsion
Ion NaCl KJ.
Metallar
Vodorod
Kovalent
Qutubsiz
Qutubli
H2 O2 N2
HCl

Kimyoviy energetika va reaksiyalarni yo’nalishi



Issiqlik effektiga ko’ra reaksiyalar

Qp
O’zgarmas bosim

Ov
O’zgarmas xajmda

Egzotermik

Endotermik

$2H_2+O_2 \rightarrow 2H_2O+O$

$N_2+O_2 \rightarrow 2NO-Q$

Termodinamikaning birinchi qonuni

$Q=U+A$

$H_2+Br_2 \rightarrow 2Hbr+68,20$ kn

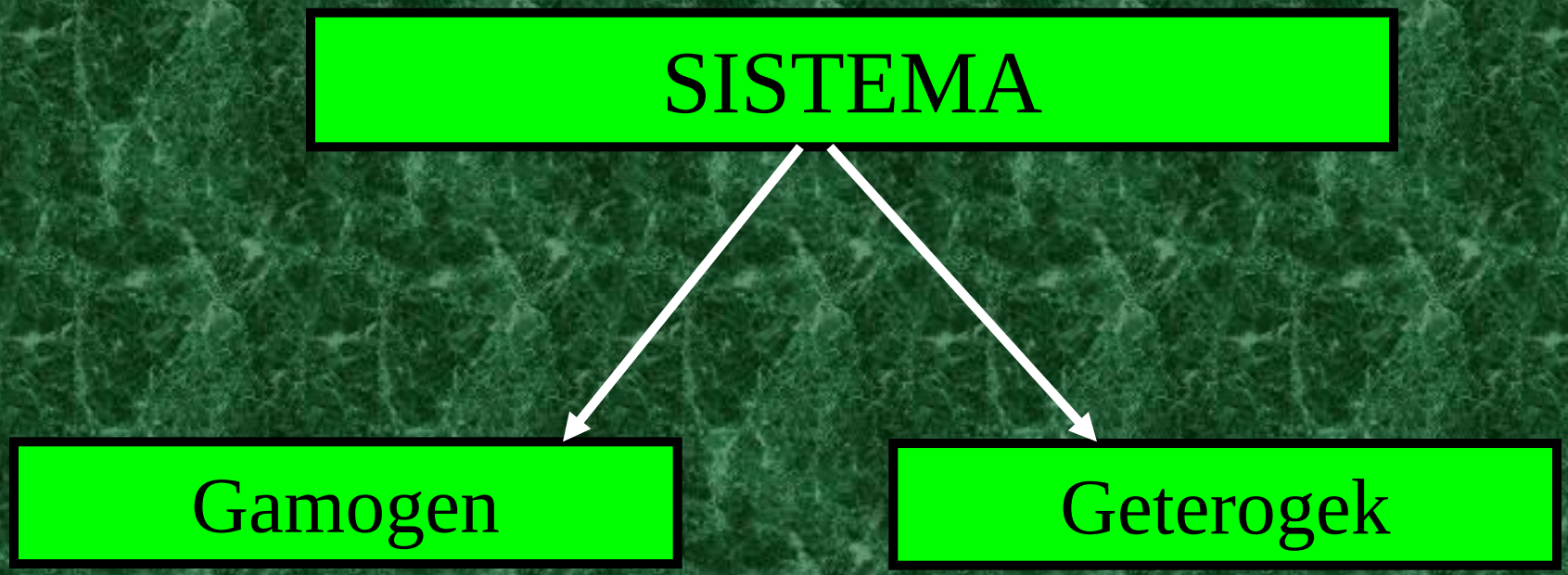
$HB_2 \rightarrow H_2+Br_2-68,20$

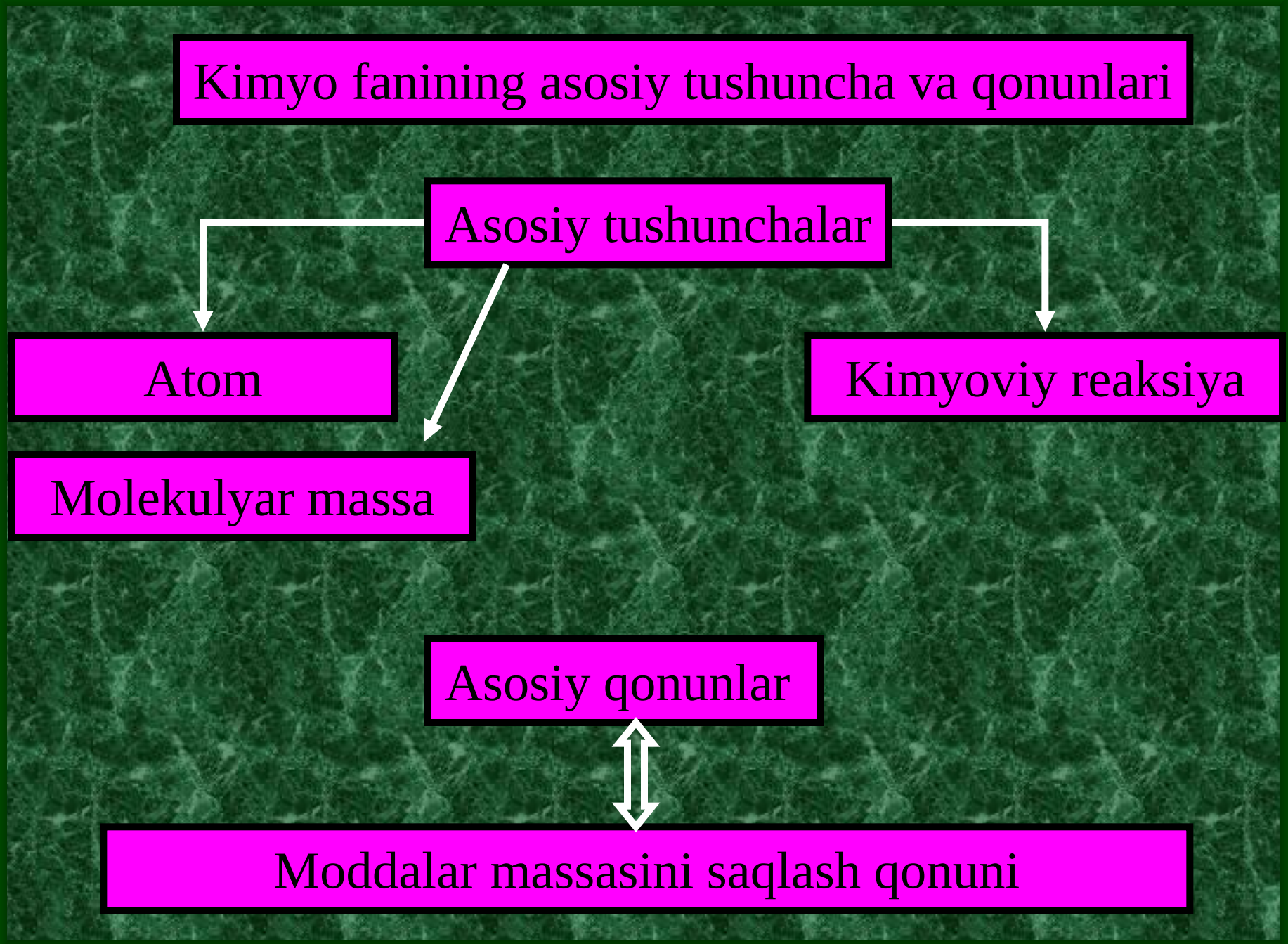
Termodinamikaning ikkinchi qonuni

$C+O_2=CO_2+412$ kj

$C+1/2\ O_2+126$ kj

$CO+1/2\ O_2=CO_2+286$ kj


SISTEMA
Gamogen
Geterogek


Kimyo fanining asosiy tushuncha va qonunlari
Asosiy tushunchalar
Atom
Molekulyar massa
Kimyoviy reaksiya
Asosiy qonunlar
Moddalar massasini saqlash qonuni

E'TIBORINGIZ UCHUN RAHMAT